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Section Experiment 30

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Determination of Iron

by Reaction with

Permanganate Titration

1. write the balanced net ionic equation for the reaction between MnO_4^- ion and Fe^{2+} ion in acidic solution.

2. How many moles of Fe^{2+} ion can be

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oxidized by $\text{IA} \times 10$
moles MnO_4^- ion in the
reaction in Question 1?
3.

Solved: Section
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Titration Section Name
son in acidic solution 1.
Write the balanced net

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ionic equation for the
reaction between MnO_4^-
ion and Fe
 $\text{MnO}_4^- + 5 \text{Fe}^{2+} + 8\text{H}^+ \rightarrow 7 \text{Mn}^{2+} + 5 \text{Fe}^{3+} + 4\text{H}_2\text{O}$

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determine the molar
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Reactions II: A Clock
Reaction? 1. A student
studied the clock
reaction described in
this experiment. she set
up Reaction Mixture 3
by mixing 10 mL 0.010
M KI, 10 mL 0.0010
Na₂S₂O₃, 20 mL 0.040
M KBrO₃, and 10 mL
M HCl. Advance Study
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Iodination of Acetone 1.

In a reaction involving

the iodination of

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acetone, the following
volumes were used to
make up the reac- tion
mixture: 5 mL. 4.0 M
acetone+10 mL. 1.0 M
HCl+10 mL.0.0050 M 1
+25 mL ...

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