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and Electron

Configurations

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Tutorial: Electron
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~~Chemistry Topic 2~~

~~Atomic structure~~

~~12.1 Electrons in~~

~~atoms HL The~~

~~Electron: Crash~~

~~Course Chemistry~~

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#5 Physics 12 |

Determination of
e/m for an electron

| Electromagnetism

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Valence Electrons
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What Is Electricity?

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Atoms | What are
They? What are

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and Electrons? How
to write electron
configurations and
what they are

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The Discovery of
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15) ~~Quantum~~

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~~9 Electrons in~~

~~atoms and the~~

~~Periodic Table~~

~~Protons Neutrons~~

~~Electrons Isotopes~~

~~Average Mass~~

~~Number \u0026amp;~~

~~Atomic Structure~~

~~Atoms vs Ions e/m~~

~~of an electron (Part~~

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13.18 Chapter 13 |

Nuclei by Ashish

Arora Chapter 13

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the outermost and

highest energy

electrons in an

atom; an atom has

eight at most. pauli

exclusion principle.

no two electrons

may have the

exact same energy

state, two

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electrons may
occupy one orbital
but they must have
opposite spins;
cannot draw two
arrows with the
same direction in
an orbital. Spin.

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Electrons move in
circular orbits .
around . the
nucleus . at . fixed.

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Electrons are never
between energy
levels or energy

shells. An electron
must have . just

the right amount of
energy. to jump

from one level to
another. A

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Chapter 4 Electrons

in Atoms. 1. How

are the wavelength

and frequency of

light related?

Wavelength. The .

amplitude. of a

wave is the wave's

height from zero to

the crest. The .

wavelength,

represented by (

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(the Greek letter lambda), is the distance between the crests. It is often measured in meters or nanometers.
Frequency

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23 – 25, 27, 31, 32,

34 – 38, 41, 45, 47,

48, 52. Section

13.1 – Models of the

Atom. Section

Review 13.1. 1. List

in chronological

order, a major

contribution of

each of these

scientists to the

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Understanding of
the atom: proposed

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Charles Page High School Dr. Stephen L. Cotton Section

13.1 Models of the

Atom OBJECTIVES:

Summarize the development of atomic theory.

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– Models of the

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Review 13.1 1. List in chronological order, a major contribution of each of these scientists to the understanding of the atom: proposed that all elements are composed of atoms. Dalton –

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particles in an
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particle duality),
light. Quantum
mechanical model.

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Atom Section
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order, a major contribution of each of these scientists to the understanding of

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electrons? To get a stable electron _____ configuration _____ (full outer shell - octet) 2. Which

group on the periodic table contains elements with stable electron configurations?

_____ noble _____

_____ gases _____

Group 18 3. If you have a full outer

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energy level you
have a stable
octet or
configuration.

4. Answers

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ompounds_Study_G
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13. Electrons In

Atoms the
outermost and
highest energy
electrons in an
atom; an atom has
eight at most. pauli
exclusion principle.
no two electrons
may have the
exact same energy
state, two
electrons may
occupy one orbital

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but they must have opposite spins; cannot draw two arrows with the same direction in an orbital.

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atoms. blank bohr

model worksheet -

blank fill in for first

20 elements ...

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fill in for first 20

elements.

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...

The key to understanding electronic arrangement is summarized in the Pauli exclusion principle: no two electrons in an

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atom can have the same set of four quantum numbers. This dramatically limits the number of electrons that can exist in a shell or a subshell.

Organization of
Electrons in Atoms

- Introductory ...

You might have
been taught that

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Atoms are the smallest particles on Earth, but inside of atoms are even smaller things called electrons. ...

The number of electrons in an atom The way atoms bond ...

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