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162 CHAPTER 6: THERMOCHEMISTRY To convert the answer to joules, we write: 101.3 J 0.18 L atm 1L atm = - · x = · w -18 J 6.17 An expansion implies an increase in volume, therefore w must be -325 J (see the defining equation for pressure-volume work.)

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164 CHAPTER 6: THERMOCHEMISTRY Substituting into the above equation: ΔE = 483.6 × 103 J - (8.314 J/mol·K)(398 K)(+1 mol) ΔE = 4.80 × 105 J = 4.80 × 102 kJ 6.28 We initially have 6 moles of gas (3 moles of chlorine and 3 moles of hydrogen). Since our product is 6

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CHAPTER 6 Thermochemistry © Houghton Mifflin Company. All rights reserved. 115 14. Two metals of equal mass with different heat capacities are subjected to the same amount of heat. Which undergoes the smallest change in temperature? a) The metal with the higher heat capacity. b) The metal with the lower heat capacity.

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Ch.6 - Thermochemistry Ch.6.1: The Nature of Energy Energy: An object's capacity to perform work or produce heat Potential Energy: Energy due to position or composition (chemical bonds). Kinetic Energy: Energy due to the motion of the object 1 2 2 KE mv Law of Conservation of Energy: Energy can neither be created nor destroyed,

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Chapter 6 - Thermochemistry - Mrs. Duffey - FHN

Chapter 6 Thermochemistry flow is determined into or out of the surroundings. Because ΔE_{univ} = 0 by the first law of thermodynamics, ΔE_{sys} = -ΔE_{sur}; what happens to the surroundings is the exact opposite of what happens to the system. chapter 6 test chemistry thermochemistry Flashcards and ... Learn chemistry test Page 6/24