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CHAPTER 2 REVIEW Measurements and Calculations  
SECTION 2 SHORT ANSWER Answer the following questions in the space provided. 1. Complete the following conversions:  
a. 100 mL 0.1 L b. 0.25 g 25 cg c. 400 cm<sup>3</sup> 0.4 L d. 400 cm<sup>3</sup> 0.0004 m 2. For each measuring device shown below, identify the quantity measured and tell when it would remain constant ...

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2(g) NO<sub>2</sub>(g) a. Balance the equation. 2N<sub>2</sub>O(g) + 3O<sub>2</sub>(g) → 4NO<sub>2</sub>(g) + 4 mol NO<sub>2</sub>:3 mol O<sub>2</sub> b. What is the mole ratio of NO<sub>2</sub> to O<sub>2</sub>? 15.0 mol c. If 20.0 mol of NO<sub>2</sub> form, how many moles of O<sub>2</sub> must have been consumed? True d. Twice as many moles of NO<sub>2</sub> form as moles of N<sub>2</sub>O are consumed. True or False? False e. Twice as many grams of NO<sub>2</sub> ...

mc06se cFMsr i-vi

CHAPTER 5 REVIEW The Periodic Law SECTION 1 SHORT ANSWER Answer the following questions in the space provided. 1. c In the modern periodic table, elements are ordered (a) according to decreasing atomic mass. (b) according to Mendeleev ' s original design. (c) according to increasing atomic number. (d) based on when they were discovered. 2. d Mendeleev noticed that certain similarities in the ...

### 5 The Periodic Law

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SECTION 2 SHORT ANSWER Answer the following questions in the space provided. 1. Use the concept of potential energy to describe how a covalent bond forms between two atoms.

### 6 Chemical Bonding

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