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Introduction to Moles ~~Mole~~

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Newton's Law of Universal Gravitation by Professor Mac  
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Chemists call that quantity the mole—a quantity of any sample whose mass is equal to its atomic mass in grams.

~~4.01 POGIL 13 Relative Mass and the Mole.pdf~~

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## ~~Relative Mass And The Mole Pogil Packet Answers~~

A mole of water molecules has a mass of 18.02 g. A mole of water molecules is  $6.022 \times 10^{23}$  molecules so 18.02 grams of water has far more molecules than 18.02 amu, which is the mass of just one molecule. A single molecule is too small to see but 18.02 grams of water is a just a bit more than a tablespoon-full.

## ~~POGIL Activity: Relative Mass and the Mole~~

a relative mass in grams  
Hydrogen  $166 \times 10^{-24}$  Carbon

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200 X 10<sup>-23</sup> Iron 930 X 10<sup>-23</sup>  
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So the mass of an average carbon atom is 12 amu and the mass of a mole of carbon atoms is 12 g. POGIL

Activity: Relative Mass and the Mole whose mass is equal to its atomic mass in grams. Relative Mass and the Mole 163.

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The Mole Answers -  
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Because 1 mole :  $6.02 \times 10^{23}$  atoms and the molar mass of an element is equal to one mole, the mass of a single atom can be calculated by dividing the mass of one mole by Avogadro's number. Arrange the following in order of mass from the smallest mass to the largest: 1.0 mol Ar,  $3.0 \times 10^{24}$  atoms Ne, 20 g Kr.

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The mole is the number of atoms which has a mass in grams equal to the numerical value of the mass of the atom in atomic mass units (amu) So the mass of an average carbon atom is 12 amu and the

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